



# Hydrogen release





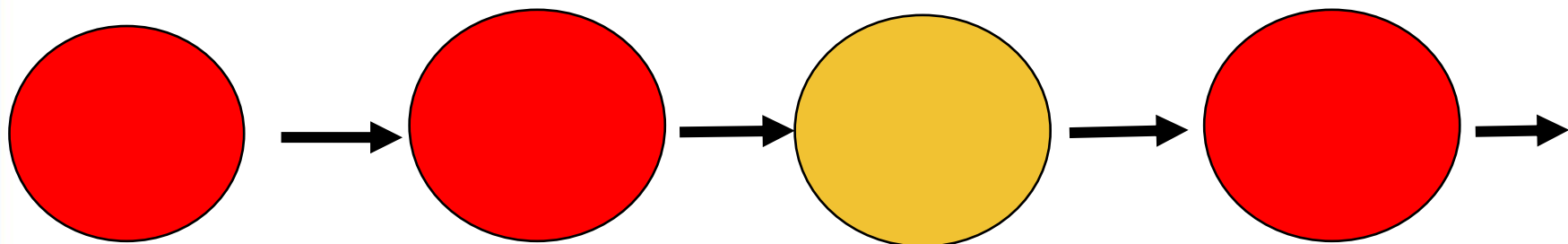
# Question

A simple method to produce gaseous hydrogen is the reaction between metal aluminum and two salts in aqueous solution (e.g. copper sulphate and sodium chloride). Investigate how the reaction rate depends on the concentration of each salt and other relevant conditions. What salts react with aluminum to release hydrogen?





# SPEAKER'S STATEMENT

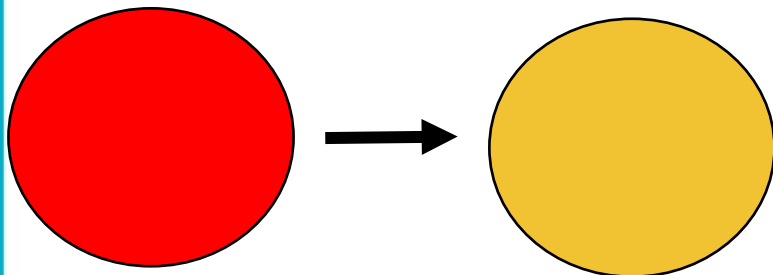


EXPLANATION

THEORY

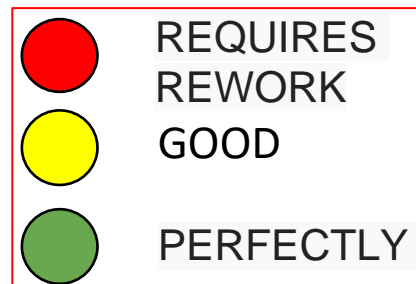
EXPERIMENT

COMPARISON



QUESTIONS  
ANSWERED

PRESENTATION  
LEVEL





# THEORY SPEAKER

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1. clear explanation of reaction

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1. do not have predictions
2. do not have mathematical model that will predict the volume of hydrogen
3. do not have theory of different type of salt
4. do not have exact explanation of oxygen film destruction





# EXPERIMENT SPEAKER

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1. understandable experiment

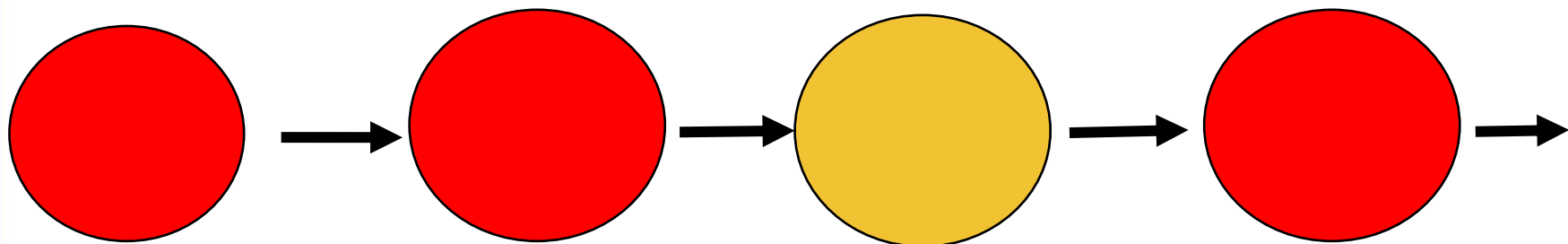
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1. did not calculate the volume of hydrogen
2. did not provide experiments with other salts
3. did not calculate the area of the surface of the aluminium
4. did not have time/volume dependence graphs





# SPEAKER'S STATEMENT

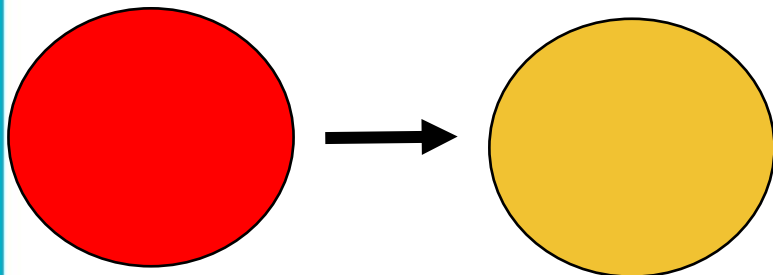


EXPLANATION

THEORY

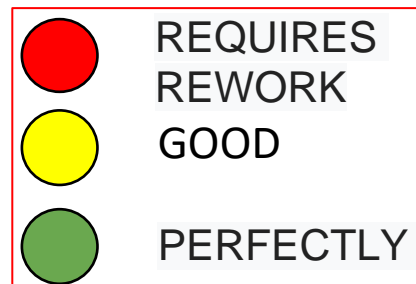
EXPERIMENT

COMPARISON



QUESTIONS  
ANSWERED

PRESENTATION  
LEVEL





# TOPICS FOR DISCUSSION

1. different types of salts.
2. volume of hydrogen.
3. surface area.







# thanks for your attention

